



ASSIGNMENT NO: 2

SUBJECT: CHEMISTRY

CLASS-XI

JULY'2026

Chapter:3 Classification of elements and periodicity in properties.

Very short answer type questions (2 marks)

1. Explain the following terms:
(a)Electronegativity (b) Electron affinity
2. Explain diagonal relationship. Give two reasons of it.
3. Define the following :(a)Enthalpy of ionisation (b) Electron gain enthalpy
4. Find IUPAC name of elements whose atomic number are 112, 104.

Short answer type questions(3 marks)

5. Give the position of following elements in periodic table: elements A & B whose atomic numbers are 25 and 20 respectively.
6. Which is strongest electronegative element and give reason for it?
7. Give reason for these:
(a)Electron affinity of 'F' is less than Cl.
(b)Enthalpy of ionisation of 'Be' is more than 'B'

Long answer type questions (5 marks)

8. Explain the following:
(a)Why fluorine show less electron affinity than chlorine?
(b)Second electron gain enthalpy of oxygen is positive.
9. s(a)Why enthalpy of ionisation Be is more than B while B is located in right of Be in same period?.
(b)Atomic radius of Ga is less than Al while Ga is located below Al in same group.
- 12.Explain the following terms(a) Shielding effect (b)Penetration effect (c) Enthalpy of atomisation..