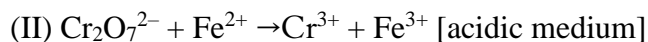
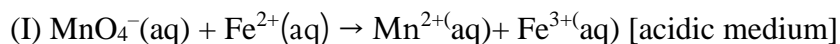


CH: Redox Reaction:

Q1. What is meant by disproportionation? Give one example.

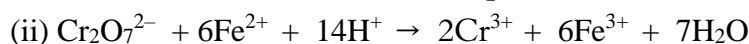
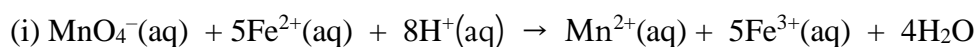
Q2. What is galvanization?

Q3. Write the balanced redox reaction .



Q4. Determine the oxidation no. of all the atoms in the following oxidants: KMnO_4 , $\text{K}_2\text{Cr}_2\text{O}_7$ and KClO_4 .

Q5. Identify the oxidizing and reducing agents in the following equations:



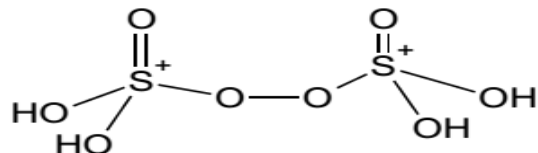
Q6. Identify the strongest and weakest reducing agent from the following metals: Zn, Cu, Na, Ag, Sn

Q7. Calculate O. S. of sulphur in the following oxoacids of 'S':



Q8. Balance the equation $\text{MnO}_4^- + \text{I}^- \rightarrow \text{Mn}^{2+} + \text{I}_2 + \text{H}_2\text{O}$ by ion electron method in acidic medium.

Q9. Indicate O.S. of each atom present in given structure of peroxodisulphuric acid.



Q10. Define:

(a) Electrochemical series (b) Redox reactions (c) Oxidizing agents

MCQ Type Questions:

. Q11.. Oxidation state of Fe in Fe_3O_4 is

(i) 5/4 (ii) 4/5 (iii) 3/2 (iv) 8/3

Q12. Oxide which cannot act as reducing agent is

(i) CO_2 (ii) ClO_2 (iii) NO_2 (iv) SO_2